

HIGH TEMPERATURE MIXING

The present invention relates to an improved process for the vapor phase catalytic oxidation of alkanes, or a mixture of alkanes and alkenes, to their 5 corresponding unsaturated carboxylic acids.

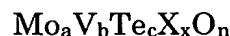
The present invention also relates to a process for the vapor phase catalytic oxidation of alkanes, or a mixture of alkanes and alkenes, in the presence of ammonia, to their corresponding unsaturated nitriles.

Moreover, the present invention relates to a process for producing an 10 improved catalyst comprising a mixed metal oxide and the catalyst so-formed.

Nitriles, such as acrylonitrile and methacrylonitrile, have been industrially produced as important intermediates for the preparation of fibers, synthetic resins, synthetic rubbers, and the like. The most popular method for producing such nitriles is to subject an olefin such as propene or isobutene to a 15 gas phase catalytic reaction with ammonia and oxygen in the presence of a catalyst at a high temperature. Known catalysts for conducting this reaction include a Mo-Bi-P-O catalyst, a V-Sb-O catalyst, an Sb-U-V-Ni-O catalyst, a Sb-Sn-O catalyst, a V-Sb-W-P-O catalyst and a catalyst obtained by mechanically mixing a V-Sb-W-O oxide and a Bi-Ce-Mo-W-O oxide. However, in view of the 20 price difference between propane and propene or between isobutane and isobutene, attention has been drawn to the development of a method for producing acrylonitrile or methacrylonitrile by an ammonoxidation reaction wherein a lower alkane, such as propane or isobutane, is used as a starting material, and it is catalytically reacted with ammonia and oxygen in a gaseous 25 phase in the presence of a catalyst.

In particular, U.S. Patent No. 5,281,745 discloses a method for producing an unsaturated nitrile comprising subjecting an alkane and ammonia in the gaseous state to catalytic oxidation in the presence of a catalyst which satisfies the conditions:

30 (1) the mixed metal oxide catalyst is represented by the empirical formula



wherein X is at least one element selected from the group consisting of niobium, tantalum, tungsten, titanium, aluminum, zirconium, chromium, manganese, iron, ruthenium, cobalt, rhodium, nickel, palladium, platinum, antimony, bismuth, boron and cerium and, when a = 1, b = 0.01 to 1.0, c = 0.01 to 1.0, x = 5 0.01 to 1.0 and n is a number such that the total valency of the metal elements is satisfied; and

(2) the catalyst has X-ray diffraction peaks at the following angles ($\pm 0.3^\circ$) of 2 θ in its X-ray diffraction pattern: 22.1°, 28.2°, 36.2°, 45.2° and 50.0°.

Similarly, Japanese Laid-Open Patent Application Publication No. 6-10 228073 discloses a method of nitrile preparation comprising reacting an alkane in a gas phase contact reaction with ammonia in the presence of a mixed metal oxide catalyst of the formula



wherein X represents one or more elements selected from niobium, tantalum, 15 titanium, aluminum, zirconium, chromium, manganese, iron, ruthenium, cobalt, rhodium, nickel, palladium, platinum, antimony, bismuth, indium and cerium and, when a = 1, b = 0.01 to 1.0, c = 0.01 to 1.0, x = 0.01 to 1.0 and n is determined by the oxide form of the elements.

Unsaturated carboxylic acids such as acrylic acid and methacrylic acid are 20 industrially important as starting materials for various synthetic resins, coating materials and plasticizers. Commercially, the current process for acrylic acid manufacture involves a two-step catalytic oxidation reaction starting with a propene feed. In the first stage, propene is converted to acrolein over a modified bismuth molybdate catalyst. In the second stage, acrolein product from the first 25 stage is converted to acrylic acid using a catalyst composed of mainly molybdenum and vanadium oxides. In most cases, the catalyst formulations are proprietary to the catalyst supplier, but, the technology is well established. Moreover, there is an incentive to develop a single step process to prepare the unsaturated acid from its corresponding alkene. Therefore, the prior art 30 describes cases where complex metal oxide catalysts are utilized for the preparation of unsaturated acid from a corresponding alkene in a single step.

Commercial incentives also exist for producing acrylic acid using a lower cost propane feed. Therefore, the prior art describes cases wherein a mixed metal oxide catalyst is used to convert propane to acrylic acid in one step.

U.S. Patent No. 5,380,933 discloses a method for producing an unsaturated carboxylic acid comprising subjecting an alkane to a vapor phase catalytic oxidation reaction in the presence of a catalyst containing a mixed metal oxide comprising, as essential components, Mo, V, Te, O and X, wherein X is at least one element selected from the group consisting of niobium, tantalum, tungsten, titanium, aluminum, zirconium, chromium, manganese, iron, ruthenium, cobalt, rhodium, nickel, palladium, platinum, antimony, bismuth, boron, indium and cerium; and wherein the proportions of the respective essential components, based on the total amount of the essential components, exclusive of oxygen, satisfy the following relationships:

0.25 < r(Mo) < 0.98, 0.003 < r(V) < 0.5, 0.003 < r(Te) < 0.5 and 0.003 < r(X) < 0.5,

wherein r(Mo), r(V), r(Te) and r(X) are the molar fractions of Mo, V, Te and X, respectively, based on the total amount of the essential components exclusive of oxygen.

U.S. Patent No. 5,994,580 discloses a process for producing acrylic acid from propane and oxygen gas through a vapor-phase catalytic oxidation reaction, the process comprising conducting the reaction using a metal oxide catalyst containing the elements Mo, V, Sb and A (wherein A is at least one element selected from the group consisting of Nb, Ta, Sn, W, Ti, Ni, Fe, Cr and Co).

Japanese Laid-Open Patent Application Publication No. 2000-037623 discloses a method for producing an unsaturated carboxylic acid comprising subjecting an alkane to a vapor phase catalytic oxidation in the presence of a catalyst having the empirical formula



wherein X is at least one element selected from the group consisting of Te and Sb, Z is at least one element selected from the group consisting of W, Cr, Ta, Ti, Zr, Hf, Mn, Re, Fe, Ru, Co, Rh, Ni, Pd, Pt, Ag, Zn, B, Al, Ga, In, Ge, Sn, Pb, P, Bi, Y, rare earth elements and alkaline earth elements, $0.1 \leq a \leq 1.0$, $0.01 \leq b \leq 1.0$,

$0.01 \leq c \leq 1.0$, $0 \leq d \leq 1.0$ and n is determined by the oxidation states of the other elements.

Despite the above-noted attempts to provide new and improved catalysts for the oxidation of alkanes to unsaturated carboxylic acids, one impediment to 5 the provision of a commercially viable process for such catalytic oxidations is the identification of a catalyst providing adequate conversion and suitable selectivity, thereby providing sufficient yield of the unsaturated product.

By the present invention, there is provided a catalyst wherein the activity and/or selectivity is enhanced and, hence, the overall yield of the desired reaction 10 product is also enhanced.

Thus, in a first embodiment, the present invention provides a process for producing an improved catalyst comprising a mixed metal oxide having the empirical formula



15 wherein N is at least one element selected from the group consisting of Te and Sb,

wherein X is at least one element selected from the group consisting of Nb, Ta, Ti, W, Se, Al, Zr, Cr, Mn, Fe, Ru, Co, Rh, Ni, Pt, Bi, B, In, As, Ge, Sn, Li, Na, K, Rb, Cs, Fr, Be, Mg, Ca, Sr, Ba, Ra, Hf, Pb, P, Pm, Eu, Gd, Dy, Ho, Er, Tm, Yb 20 and Lu,

wherein Z is at least one element selected from the group consisting of Au, Ag, Re, Pr, Zn, Ga, Pd, Ir, Nd, Y, Sm, Tb, Br, Cu, Sc, Cl, F and I, and

wherein, when $a = 1$, $b = 0.01$ to 1.0 , $c = 0.01$ to 1.0 , $d = 0.01$ to 1.0 , $e = 0$ to 0.1 and f is dependent on the oxidation state of the other elements, comprising:

25 (a) admixing compounds of elements Mo, V, N, X and Z and at least one solvent at a temperature greater than or equal to 45°C to form an admixture,

wherein Mo, V, N, X and Z are present in such amounts that the atomic ratio of Mo : V : N : X : Z is $a : b : c : d : e$, and, when $a = 1$, $b = 0.01$ to 1.0 , $c = 0.1$ to 1.0 , $d = 0.01$ to 1.0 and $e = 0$ to 0.1 ;

(b) removing said at least one solvent from the so-formed admixture to form a catalyst precursor; and

(c) calcining said catalyst precursor to obtain said mixed metal oxide.

In a second embodiment, the present invention provides an improved catalyst prepared by the process of the first embodiment.

In a third embodiment, the present invention provides a process for

5 producing an unsaturated carboxylic acid which comprises subjecting an alkane, or a mixture of an alkane and an alkene, to a vapor phase catalytic oxidation reaction in the presence of a catalyst comprising a mixed metal oxide having the empirical formula



10 wherein N is at least one element selected from the group consisting of Te and Sb,

wherein X is at least one element selected from the group consisting of Nb, Ta, Ti, W, Se, Al, Zr, Cr, Mn, Fe, Ru, Co, Rh, Ni, Pt, Bi, B, In, As, Ge, Sn, Li, Na, K, Rb, Cs, Fr, Be, Mg, Ca, Sr, Ba, Ra, Hf, Pb, P, Pm, Eu, Gd, Dy, Ho, Er, Tm, Yb and Lu,

15 wherein Z is at least one element selected from the group consisting of Au, Ag, Re, Pr, Zn, Ga, Pd, Ir, Nd, Y, Sm, Tb, Br, Cu, Sc, Cl, F and I, and

wherein, when a = 1, b = 0.01 to 1.0, c = 0.01 to 1.0, d = 0.01 to 1.0, e = 0 to 0.1 and f is dependent on the oxidation state of the other elements, prepared by the

20 process comprising:

(a) admixing compounds of elements Mo, V, N, X and Z and at least one solvent at a temperature greater than or equal to 45°C to form an admixture,

25 wherein Mo, V, N, X and Z are present in such amounts that the atomic ratio of Mo : V : N : X : Z is a : b : c : d : e, and, when a = 1, b = 0.01 to 1.0, c = 0.1 to 1.0, d = 0.01 to 1.0 and e = 0 to 0.1;

(b) removing said at least one solvent from the so-formed admixture to form a catalyst precursor; and

(c) calcining said catalyst precursor to obtain said mixed metal oxide.

30 In a fourth embodiment, the present invention provides a process for producing an unsaturated nitrile which comprises subjecting an alkane, or a mixture of an alkane and an alkene, and ammonia to a vapor phase catalytic

oxidation reaction in the presence of a catalyst comprising a mixed metal oxide having the empirical formula



wherein N is at least one element selected from the group consisting of Te and

5 Sb,

wherein X is at least one element selected from the group consisting of Nb, Ta, Ti, W, Se, Al, Zr, Cr, Mn, Fe, Ru, Co, Rh, Ni, Pt, Bi, B, In, As, Ge, Sn, Li, Na, K, Rb, Cs, Fr, Be, Mg, Ca, Sr, Ba, Hf, Pb, P, Pm, Eu, Gd, Dy, Ho, Er, Tm, Yb and Lu,

10 wherein Z is at least one element selected from the group consisting of Au, Ag, Re, Pr, Zn, Ga, Pd, Ir, Nd, Y, Sm, Tb, Br, Cu, Sc, Cl, F and I, and wherein, when a = 1, b = 0.01 to 1.0, c = 0.01 to 1.0, d = 0.01 to 1.0, e = 0 to 0.1 and f is dependent on the oxidation state of the other elements, prepared by the process comprising:

15 (a) admixing compounds of elements Mo, V, N, X and Z and at least one solvent at a temperature greater than or equal to 45°C to form an admixture,

wherein Mo, V, N, X and Z are present in such amounts that the atomic ratio of Mo : V : N : X : Z is a : b : c : d : e, and, when a = 1, b = 0.01 to 1.0, c = 0.1 to 1.0, d = 0.01 to 1.0 and e = 0 to 0.1;

20 (b) removing said at least one solvent from the so-formed admixture to form a catalyst precursor; and

(c) calcining said catalyst precursor to obtain said mixed metal oxide.

25 The mixed metal oxide to be used as a catalyst component of the present invention has the empirical formula



wherein N is at least one element selected from the group consisting of Te and Sb,

30 wherein X is at least one element selected from the group consisting of Nb, Ta, Ti, W, Se, Al, Zr, Cr, Mn, Fe, Ru, Co, Rh, Ni, Pt, Bi, B, In, As, Ge, Sn, Li, Na, K, Rb, Cs, Fr, Be, Mg, Ca, Sr, Ba, Hf, Pb, P, Pm, Eu, Gd, Dy, Ho, Er, Tm, Yb and Lu,

wherein Z is at least one element selected from the group consisting of Au, Ag, Re, Pr, Zn, Ga, Pd, Ir, Nd, Y, Sm, Tb, Br, Cu, Sc, Cl, F and I, and wherein, when a = 1, b = 0.01 to 1.0, c = 0.01 to 1.0, d = 0.01 to 1.0, e = 0 to 0.1 and f is dependent on the oxidation state of the other elements.

5 Preferably, when a = 1, b = 0.1 to 0.5, c = 0.05 to 0.5, d = 0.01 to 0.5 and e = 0.001 to 0.1. More preferably, when a = 1, b = 0.15 to 0.45, c = 0.05 to 0.45, d = 0.01 to 0.1 and e = 0.001 to 0.1. The value of f, i.e. the amount of oxygen present, is dependent on the oxidation state of the other elements in the catalyst. However, f is typically in the range of from 3 to 4.7.

10 Preferred promoted mixed metal oxides have the empirical formulae $Mo_aV_bTe_cNb_dZe_f$ and $Mo_aV_bSb_cNb_dZe_f$ wherein Z, a, b, c, d, e and f are as previously defined.

The mixed metal oxide can be prepared in the following manner.

15 In a first step an admixture, i.e. a slurry or a solution, is formed by admixing metal compounds, preferably at least one of which contains oxygen, and at least one solvent, at a temperature greater than or equal to 45°C, preferably greater than or equal to 50°C, more preferably greater than or equal to 60°C, in appropriate amounts to form the slurry or solution. The metal compounds contain elements Mo, V, N, X and Z, as previously defined. The 20 admixing of the metal compounds and the at least one solvent at a temperature greater than or equal to 45°C means that upon completion of the addition of all the components to the admixture, the admixture is at a temperature of greater than or equal to 45°C. For example, each ingredient added to the admixture may be at a temperature greater than or equal to 45°C. Alternatively, each ingredient 25 may be at a temperature greater than or less than 60°C so long as, when all of the ingredients are added together, the resultant temperature of the admixture is greater than or equal to 60°C. There is no particular upper limit to the temperature at which the admixture is formed, but, for ease of handling, 100°C can be used as a practicable limit.

30 Suitable solvents include water; alcohols including, but not limited to, methanol, ethanol, propanol, and diols, etc.; as well as other polar solvents known in the art. Generally, water is preferred. The water is any water suitable

for use in chemical syntheses including, without limitation, distilled water and de-ionized water. The amount of water present is preferably an amount sufficient to keep the elements substantially in solution long enough to avoid or minimize compositional and/or phase segregation during the preparation steps.

5 Accordingly, the amount of water will vary according to the amounts and solubilities of the materials combined.

For example, when a mixed metal oxide of the formula $Mo_aV_bTe_cNb_dO_f$, wherein the element N is Te and the element X is Nb, is to be prepared, an aqueous solution of niobium oxalate may be added to an aqueous solution or 10 slurry of ammonium heptamolybdate, ammonium metavanadate and telluric acid, so that the atomic ratio of the respective metal elements would be in the prescribed proportions.

Once the admixture, i.e. the aqueous slurry or solution, is formed, the water is removed by any suitable method, known in the art, to form a catalyst 15 precursor. Such methods include, without limitation, vacuum drying, freeze drying, spray drying, rotary evaporation and air drying. Vacuum drying is generally performed at pressures ranging from 10mmHg to 500mmHg. Freeze drying typically entails freezing the slurry or solution, using, for instance, liquid nitrogen, and drying the frozen slurry or solution under vacuum. Spray drying is 20 generally performed under an inert atmosphere such as nitrogen or argon, with an inlet temperature ranging from 125°C to 200°C and an outlet temperature ranging from 75°C to 150°C. Rotary evaporation is generally performed at a bath temperature of from 25°C to 90°C and at a pressure of from 10mmHg to 760mmHg, preferably at a bath temperature of from 40° to 90°C and at a 25 pressure of from 10mmHg to 350mmHg, more preferably at a bath temperature of from 40°C to 60°C and at a pressure of from 10mmHg to 40mmHg. Air drying may be effected at temperatures ranging from 25°C to 90°C. Rotary evaporation is generally utilized.

Once obtained, the catalyst precursor is calcined. The calcination may be 30 conducted in an oxygen-containing atmosphere or in the substantial absence of oxygen, e.g., in an inert atmosphere or in vacuo. The inert atmosphere may be any material which is substantially inert, i.e., does not react or interact with, the

catalyst precursor. Suitable examples include, without limitation, nitrogen, argon, xenon, helium or mixtures thereof. Preferably, the inert atmosphere is argon or nitrogen. The inert atmosphere may flow over the surface of the catalyst precursor or may not flow thereover (a static environment). When the inert 5 atmosphere does flow over the surface of the catalyst precursor, the flow rate can vary over a wide range, e.g., at a space velocity of from 1 to 500 hr⁻¹.

The calcination is usually performed at a temperature of from 350°C to 850°C, preferably from 400°C to 700°C, more preferably from 500°C to 640°C. The calcination is performed for an amount of time suitable to form the 10 aforementioned catalyst. Typically, the calcination is performed for from 0.5 to 30 hours, preferably from 1 to 25 hours, more preferably for from 1 to 15 hours, to obtain the desired promoted mixed metal oxide.

In a preferred mode of operation, the catalyst precursor is calcined in two stages. In the first stage, the catalyst precursor is calcined in an oxidizing 15 environment (e.g. air), e.g., at a temperature of from 200°C to 400°C, preferably from 275°C to 325°C for from 15 minutes to 8 hours, preferably for from 1 to 3 hours. In the second stage, the material from the first stage is calcined in a non-oxidizing environment (e.g., an inert atmosphere), e.g., at a temperature of from 500°C to 750°C, preferably for from 550°C to 650°C, for 15 minutes to 8 hours, 20 preferably for from 1 to 3 hours. Optionally, a reducing gas, such as, for example, ammonia or hydrogen, may be added during the second stage calcination.

In a particularly preferred mode of operation, the catalyst precursor in the first stage is placed in the desired oxidizing atmosphere at room temperature and then raised to the first stage calcination temperature and held there for the 25 desired first stage calcination time. The atmosphere is then replaced with the desired non-oxidizing atmosphere for the second stage calcination, the temperature is raised to the desired second stage calcination temperature and held there for the desired second stage calcination time.

Although any type of heating mechanism, e.g., a furnace, may be utilized 30 during the calcination, it is preferred to conduct the calcination under a flow of the designated gaseous environment. Therefore, it is advantageous to conduct

the calcination in a bed with continuous flow of the desired gas(es) through the bed of solid catalyst precursor particles.

With calcination, a catalyst is formed having the formula $Mo_aV_bN_cX_dZ_eOf$ wherein N, X, Z, a, b, c, d, e and f are as previously defined.

5 The starting materials for the above promoted mixed metal oxide are not limited to those described above. A wide range of materials including, for example, oxides, nitrates, halides or oxyhalides, alkoxides, acetylacetones, and organometallic compounds may be used. For example, ammonium heptamolybdate may be utilized for the source of molybdenum in the catalyst.

10 However, compounds such as MoO_3 , MoO_2 , $MoCl_5$, $MoOCl_4$, $Mo(OC_2H_5)_5$, molybdenum acetylacetone, phosphomolybdc acid and silicomolybdc acid may also be utilized instead of ammonium heptamolybdate. Similarly, ammonium metavanadate may be utilized for the source of vanadium in the catalyst. However, compounds such as V_2O_5 , V_2O_3 , $VOCl_3$, VCl_4 , $VO(OC_2H_5)_3$, vanadyl sulfate, vanadium acetylacetone and vanadyl acetylacetone may also be utilized instead of ammonium metavanadate. The tellurium source may include telluric acid, $TeCl_4$, $Te(OC_2H_5)_5$, $Te(OCH(CH_3)_2)_4$ and TeO_2 . The niobium source may include ammonium niobium oxalate, Nb_2O_5 , $NbCl_5$, niobic acid or $Nb(OC_2H_5)_5$ as well as the more conventional niobium oxalate.

20 A mixed metal oxide, thus obtained, exhibits excellent catalytic activities by itself. However, the mixed metal oxide may be converted to a catalyst having higher activities by contact with a liquid contact member selected from the group consisting of organic acids, alcohols, inorganic acids and hydrogen peroxide.

25 Contacting with a liquid contact member selected from the group consisting of organic acids, alcohols, inorganic acids and hydrogen peroxide may be effected without any particular restrictions. In this regard, the liquid contact member is normally used in an amount of 1 to 100 times the volume of the mixed metal oxide or the first calcined precursor, preferably 3 to 50 times the volume, more preferably 5 to 25 times the volume. Contacting at elevated temperatures 30 is permissible, however, if prolonged contact time is not a consideration, contacting at room temperature may be utilized. Normally, contact temperatures of room temperature to 100°C are utilized, preferably 50°C to 90°C, more

preferably 60°C to 80°C. As previously noted, contact time will be affected by the temperature at which the contacting is carried out. Normally, contact times of 1 to 100 hours are utilized, preferably 2 to 20 hours, more preferably 5 to 10 hours. The contact mixture is preferably agitated during the contacting.

5 There are no particular restrictions upon the organic acids which may be used as the liquid contacting member. For example, oxalic acid, formic acid, acetic acid, citric acid and tartaric acid may be used, however, oxalic acid is preferred. If the organic acid is a liquid, it may be used as is or in an aqueous solution. If the organic acid is a solid, it is used in an aqueous solution. When 10 using aqueous solutions, there are no particular restrictions on the concentration of the organic acid. Normally, the concentration of the organic acid in the aqueous solution can vary from 0.1 to 50% by weight, preferably 1 to 15 % by weight.

15 There are no particular restrictions upon the alcohols which may be used as the liquid contacting member. For example, methanol, ethanol, propanol, butanol, hexanol and diols may be utilized, however, alcohols having one to four carbon atoms are preferred, with ethylene glycol being particularly preferred. The alcohols may be utilized in the form of aqueous solutions, but, if so, the water content should be held to 20 % by weight or less for the best effectiveness.

20 Similarly, there are no particular restrictions upon the inorganic acids which may be used as the liquid contacting member. For example, telluric acid, nitric acid, sulfuric acid, phosphoric acid, hydrochloric acid, perchloric acid, chloric acid and hypochlorous acid may be used. The inorganic acids are typically used as aqueous solutions with concentrations of the acids in the range of from 25 0.1 to 50 % by weight, preferably from 0.1 to 10 % by weight.

When hydrogen peroxide is utilized as the liquid contacting member, it is used in the form of an aqueous solution having a concentration in the range of from 0.1 to 50 % by weight, preferably from 1 to 10% by weight.

30 After contacting with the liquid contacting member, insoluble material is recovered from the so-formed contact mixture. The insoluble material may be recovered by any conventional method, e.g., centrifugation or filtration. If the contacting was conducted at elevated temperature, the contact mixture may be

cooled prior to recovery of the insoluble material. If desired, the recovered insoluble material may be washed with one of the previously disclosed solvents, preferably water. Moreover, the recovered insoluble material may be recalcined. If such recalcination is effected, it is preferred that it be carried out under an 5 inert atmosphere.

Additionally, the mixed metal oxide may be converted to a catalyst having higher activities by grinding.

There is no particular restriction as to the grinding method, and conventional methods may be employed. As a dry grinding method, a method of 10 using a gas stream grinder may, for example, be mentioned wherein coarse particles are permitted to collide with one another in a high speed gas stream for grinding. The grinding may be conducted not only mechanically but also by using a mortar or the like in the case of a small scale operation.

As a wet grinding method wherein grinding is conducted in a wet state by 15 adding water or an organic solvent to the above mixed metal oxide, a conventional method of using a rotary cylinder-type medium mill or a medium-stirring type mill, may be mentioned. The rotary cylinder-type medium mill is a wet mill of the type wherein a container for the object to be ground is rotated, and it includes, for example, a ball mill and a rod mill. The medium-stirring type 20 mill is a wet mill of the type wherein the object to be ground, contained in a container is stirred by a stirring apparatus, and it includes, for example, a rotary screw type mill, and a rotary disc type mill.

The conditions for grinding may suitably be set to meet the nature of the above-mentioned promoted mixed metal oxide, the viscosity, the concentration, 25 etc. of the solvent used in the case of wet grinding, or the optimum conditions of the grinding apparatus.

Further, in some cases, it is possible to further improve the catalytic activities by further adding a solvent to the ground catalyst precursor to form a solution or slurry, followed by drying again. There is no particular restriction as 30 to the concentration of the solution or slurry, and it is usual to adjust the solution or slurry so that the total amount of the starting material compounds for the ground catalyst precursor is from 10 to 60 wt %. Then, this solution or

slurry is dried by a method such as spray drying, freeze drying, evaporation to dryness or vacuum drying, preferably by the spray drying method. Further, similar drying may be conducted also in the case where wet grinding is conducted.

5 The oxide obtained by the above-mentioned method may be used as a final catalyst, but it may further be subjected to heat treatment usually at a temperature of from 200° to 700°C for from 0.1 to 10 hours.

10 The mixed metal oxide thus obtained may be used by itself as a solid catalyst, but may be formed into a catalyst together with a suitable carrier such as silica, alumina, titania, aluminosilicate, diatomaceous earth or zirconia. Further, it may be molded into a suitable shape and particle size depending upon the scale or system of the reactor.

15 Alternatively, the metal components of the presently contemplated catalyst may be supported on materials such as alumina, silica, silica-alumina, zirconia, titania, etc. by conventional incipient wetness techniques. In one typical method, solutions containing the metals are contacted with the dry support such that the support is wetted; then, the resultant wetted material is dried, for example, at a temperature from room temperature to 200°C followed by calcination as described above. In another method, metal solutions are contacted 20 with the support, typically in volume ratios of greater than 3 : 1 (metal solution : support), and the solution agitated such that the metal ions are ion-exchanged onto the support. The metal-containing support is then dried and calcined as detailed above. In any event, the admixture of the components, i.e. metal compounds and solvent(s), must produce an admixture having a temperature 25 greater than or equal to 45°C.

30 As previously noted, the present invention provides a process for producing an unsaturated carboxylic acid, which comprises subjecting an alkane, or a mixture of an alkane and an alkene, to a vapor phase catalytic oxidation reaction in the presence of the aforementioned catalyst to produce an unsaturated carboxylic acid.

In the production of such an unsaturated carboxylic acid, it is preferred to employ a starting material gas which contains steam. In such a case, as a

starting material gas to be supplied to the reaction system, a gas mixture comprising a steam-containing alkane, or a steam-containing mixture of alkane and alkene, and an oxygen-containing gas, is usually used. However, the steam-containing alkane, or the steam-containing mixture of alkane and alkene, and the oxygen-containing gas may be alternately supplied to the reaction system. The steam to be employed may be present in the form of steam gas in the reaction system, and the manner of its introduction is not particularly limited.

Further, as a diluting gas, an inert gas such as nitrogen, argon or helium may be supplied. The molar ratio (alkane or mixture of alkane and alkene) : (oxygen) : (diluting gas) : (H₂O) in the starting material gas is preferably (1) : (0.1 to 10) : (0 to 20) : (0.2 to 70), more preferably (1) : (1 to 5.0) : (0 to 10) : (5 to 40).

When steam is supplied together with the alkane, or the mixture of alkane and alkene, as starting material gas, the selectivity for an unsaturated carboxylic acid is distinctly improved, and the unsaturated carboxylic acid can be obtained from the alkane, or mixture of alkane and alkene, in good yield simply by contacting in one stage. However, the conventional technique utilizes a diluting gas such as nitrogen, argon or helium for the purpose of diluting the starting material. As such a diluting gas, to adjust the space velocity, the oxygen partial pressure and the steam partial pressure, an inert gas such as nitrogen, argon or helium may be used together with the steam.

As the starting material alkane it is preferred to employ a C₃-alkane, particularly propane, isobutane or n-butane; more preferably, propane or isobutane; most preferably, propane. According to the present invention, from such an alkane, an unsaturated carboxylic acid such as an α,β-unsaturated carboxylic acid can be obtained in good yield. For example, when propane or isobutane is used as the starting material alkane, acrylic acid or methacrylic acid will be obtained, respectively, in good yield.

In the present invention, as the starting material mixture of alkane and alkene, it is preferred to employ a mixture of C₃-alkane and C₃-alkene, particularly propane and propene, isobutane and isobutene or n-butane and n-butene. As the starting material mixture of alkane and alkene, propane and

propene or isobutane and isobutene are more preferred. Most preferred is a mixture of propane and propene. According to the present invention, from such a mixture of an alkane and an alkene, an unsaturated carboxylic acid such as an α,β -unsaturated carboxylic acid can be obtained in good yield. For example,

5 when propane and propene or isobutane and isobutene are used as the starting material mixture of alkane and alkene, acrylic acid or methacrylic acid will be obtained, respectively, in good yield. Preferably, in the mixture of alkane and alkene, the alkene is present in an amount of at least 0.5% by weight, more preferably at least 1.0% by weight to 95% by weight; most preferably, 3% by

10 weight to 90% by weight.

As an alternative, an alkanol, such as isobutanol, which will dehydrate under the reaction conditions to form its corresponding alkene, i.e. isobutene, may also be used as a feed to the present process or in conjunction with the previously mentioned feed streams.

15 The purity of the starting material alkane is not particularly limited, and an alkane containing a lower alkane such as methane or ethane, air or carbon dioxide, as impurities, may be used without any particular problem. Further, the starting material alkane may be a mixture of various alkanes. Similarly, the purity of the starting material mixture of alkane and alkene is not particularly limited, and a mixture of alkane and alkene containing a lower alkene such as ethene, a lower alkane such as methane or ethane, air or carbon dioxide, as impurities, may be used without any particular problem. Further, the starting material mixture of alkane and alkene may be a mixture of various alkanes and alkenes.

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25 There is no limitation on the source of the alkene. It may be purchased, per se, or in admixture with an alkane and/or other impurities. Alternatively, it can be obtained as a by-product of alkane oxidation. Similarly, there is no limitation on the source of the alkane. It may be purchased, per se, or in admixture with an alkene and/or other impurities. Moreover, the alkane, 30 regardless of source, and the alkene, regardless of source, may be blended as desired.

The detailed mechanism of the oxidation reaction of the present invention is not clearly understood, but the oxidation reaction is carried out by oxygen atoms present in the above promoted mixed metal oxide or by molecular oxygen present in the feed gas. To incorporate molecular oxygen into the feed gas, such 5 molecular oxygen may be pure oxygen gas. However, it is usually more economical to use an oxygen-containing gas such as air, since purity is not particularly required.

It is also possible to use only an alkane, or a mixture of alkane and alkene, substantially in the absence of molecular oxygen for the vapor phase catalytic 10 reaction. In such a case, it is preferred to adopt a method wherein a part of the catalyst is appropriately withdrawn from the reaction zone from time to time, then sent to an oxidation regenerator, regenerated and then returned to the reaction zone for reuse. As the regeneration method of the catalyst, a method may, for example, be mentioned which comprises contacting an oxidative gas 15 such as oxygen, air or nitrogen monoxide with the catalyst in the regenerator usually at a temperature of from 300° to 600°C.

The process of the present invention will be described in further detail with respect to a case where propane is used as the starting material alkane and air is used as the oxygen source. The reaction system may be a fixed bed system 20 or a fluidized bed system. However, since the reaction is an exothermic reaction, a fluidized bed system may preferably be employed whereby it is easy to control the reaction temperature. The proportion of air to be supplied to the reaction system is important for the selectivity for the resulting acrylic acid, and it is usually at most 25 moles, preferably from 0.2 to 18 moles per mole of propane, 25 whereby high selectivity for acrylic acid can be obtained. This reaction can be conducted usually under atmospheric pressure, but may be conducted under a slightly elevated pressure or slightly reduced pressure. With respect to other alkanes such as isobutane, or to mixtures of alkanes and alkenes such as propane and propene, the composition of the feed gas may be selected in 30 accordance with the conditions for propane.

Typical reaction conditions for the oxidation of propane or isobutane to acrylic acid or methacrylic acid may be utilized in the practice of the present

invention. The process may be practiced in a single pass mode (only fresh feed is fed to the reactor) or in a recycle mode (at least a portion of the reactor effluent is returned to the reactor). General conditions for the process of the present invention are as follows: the reaction temperature can vary from 200°C to 700°C, 5 but is usually in the range of from 200°C to 550°C, more preferably 250°C to 480°C, most preferably 300°C to 400°C; the gas space velocity, SV, in the vapor phase reaction is usually within a range of from 100 to 10,000 hr⁻¹, preferably 300 to 6,000 hr⁻¹, more preferably 300 to 2,000 hr⁻¹; the average contact time 10 with the catalyst can be from 0.01 to 10 seconds or more, but is usually in the range of from 0.1 to 10 seconds, preferably from 2 to 6 seconds; the pressure in the reaction zone usually ranges from 0 to 75 psig, but is preferably no more than 50 psig. In a single pass mode process, it is preferred that the oxygen be supplied from an oxygen-containing gas such as air. The single pass mode 15 process may also be practiced with oxygen addition. In the practice of the recycle mode process, oxygen gas by itself is the preferred source so as to avoid the build up of inert gases in the reaction zone.

Of course, in the oxidation reaction of the present invention, it is important that the hydrocarbon and oxygen concentrations in the feed gases be maintained at the appropriate levels to minimize or avoid entering a flammable 20 regime within the reaction zone or especially at the outlet of the reactor zone. Generally, it is preferred that the outlet oxygen levels be low to both minimize after-burning and, particularly, in the recycle mode of operation, to minimize the amount of oxygen in the recycled gaseous effluent stream. In addition, operation 25 of the reaction at a low temperature (below 450°C) is extremely attractive because after-burning becomes less of a problem which enables the attainment of higher selectivity to the desired products. The catalyst of the present invention operates more efficiently at the lower temperature range set forth above, significantly reducing the formation of acetic acid and carbon oxides, and increasing selectivity to acrylic acid. As a diluting gas to adjust the space velocity 30 and the oxygen partial pressure, an inert gas such as nitrogen, argon or helium may be employed.

When the oxidation reaction of propane, and especially the oxidation reaction of propane and propene, is conducted by the method of the present invention, carbon monoxide, carbon dioxide, acetic acid, etc. may be produced as by-products, in addition to acrylic acid. Further, in the method of the present invention, an unsaturated aldehyde may sometimes be formed depending upon the reaction conditions. For example, when propane is present in the starting material mixture, acrolein may be formed; and when isobutane is present in the starting material mixture, methacrolein may be formed. In such a case, such an unsaturated aldehyde can be converted to the desired unsaturated carboxylic acid by subjecting it again to the vapor phase catalytic oxidation with the promoted mixed metal oxide-containing catalyst of the present invention or by subjecting it to a vapor phase catalytic oxidation reaction with a conventional oxidation reaction catalyst for an unsaturated aldehyde.

As previously noted, the present invention also provides a process for producing an unsaturated nitrile, which comprises subjecting an alkane, or a mixture of an alkane and an alkene, and ammonia to a vapor phase catalytic oxidation reaction in the presence of the aforementioned catalyst to produce an unsaturated nitrile.

In the production of such an unsaturated nitrile, as the starting material alkane, it is preferred to employ a C₃-8 alkane such as propane, butane, isobutane, pentane, hexane and heptane. However, in view of the industrial application of nitriles to be produced, it is preferred to employ a lower alkane having 3 or 4 carbon atoms, particularly propane and isobutane.

Similarly, as the starting material mixture of alkane and alkene, it is possible to employ a mixture of C₃-8 alkane and C₃-8 alkene such as propane and propene, butane and butene, isobutane and isobutene, pentane and pentene, hexane and hexene, and heptane and heptene. However, in view of the industrial application of nitriles to be produced, it is more preferred to employ a mixture of a lower alkane having 3 or 4 carbon atoms and a lower alkene having 3 or 4 carbon atoms, particularly propane and propene or isobutane and isobutene. Preferably, in the mixture of alkane and alkene, the alkene is present in an

amount of at least 0.5% by weight, more preferably at least 1.0% by weight to 95% by weight, most preferably 3% by weight to 90% by weight.

The purity of the starting material alkane is not particularly limited, and an alkane containing a lower alkane such as methane or ethane, air or carbon dioxide, as impurities, may be used without any particular problem. Further, the starting material alkane may be a mixture of various alkanes. Similarly, the purity of the starting material mixture of alkane and alkene is not particularly limited, and a mixture of alkane and alkene containing a lower alkene such as ethene, a lower alkane such as methane or ethane, air or carbon dioxide, as impurities, may be used without any particular problem. Further, the starting material mixture of alkane and alkene may be a mixture of various alkanes and alkenes.

There is no limitation on the source of the alkene. It may be purchased, per se, or in admixture with an alkane and/or other impurities. Alternatively, it can be obtained as a by-product of alkane oxidation. Similarly, there is no limitation on the source of the alkane. It may be purchased, per se, or in admixture with an alkene and/or other impurities. Moreover, the alkane, regardless of source, and the alkene, regardless of source, may be blended as desired.

The detailed mechanism of the ammonoxidation reaction of this aspect of the present invention is not clearly understood. However, the oxidation reaction is conducted by the oxygen atoms present in the above promoted mixed metal oxide or by the molecular oxygen in the feed gas. When molecular oxygen is incorporated in the feed gas, the oxygen may be pure oxygen gas. However, since high purity is not required, it is usually economical to use an oxygen-containing gas such as air.

As the feed gas, it is possible to use a gas mixture comprising an alkane, or a mixture of an alkane and an alkene, ammonia and an oxygen-containing gas. However, a gas mixture comprising an alkane or a mixture of an alkane and an alkene and ammonia, and an oxygen-containing gas may be supplied alternately.

When the gas phase catalytic reaction is conducted using an alkane, or a mixture of an alkane and an alkene, and ammonia substantially free from molecular oxygen, as the feed gas, it is advisable to employ a method wherein a part of the catalyst is periodically withdrawn and sent to an oxidation

5 regenerator for regeneration, and the regenerated catalyst is returned to the reaction zone. As a method for regenerating the catalyst, a method may be mentioned wherein an oxidizing gas such as oxygen, air or nitrogen monoxide is permitted to flow through the catalyst in the regenerator usually at a temperature of from 300°C to 600°C.

10 These aspects of the present invention will be described in further detail with respect to a case where propane is used as the starting material alkane and air is used as the oxygen source. The proportion of air to be supplied for the reaction is important with respect to the selectivity for the resulting acrylonitrile. Namely, high selectivity for acrylonitrile is obtained when air is

15 supplied within a range of at most 25 moles, particularly 1 to 15 moles, per mole of the propane. The proportion of ammonia to be supplied for the reaction is preferably within a range of from 0.2 to 5 moles, particularly from 0.5 to 3 moles, per mole of propane. This reaction may usually be conducted under atmospheric pressure, but may be conducted under a slightly increased pressure or a slightly

20 reduced pressure. With respect to other alkanes such as isobutane, or to mixtures of alkanes and alkenes such as propane and propene, the composition of the feed gas may be selected in accordance with the conditions for propane.

The processes of these aspects of the present invention may be conducted at a temperature of, for example, from 250°C to 480°C. More preferably, the

25 temperature is from 300°C to 400°C. The gas space velocity, SV, in the gas phase reaction is usually within the range of from 100 to 10,000 hr⁻¹, preferably from 300 to 6,000 hr⁻¹, more preferably from 300 to 2,000 hr⁻¹. As a diluent gas, for adjusting the space velocity and the oxygen partial pressure, an inert gas such as nitrogen, argon or helium can be employed. When ammoxidation of propane is

30 conducted by the method of the present invention, in addition to acrylonitrile, carbon monoxide, carbon dioxide, acetonitrile, hydrocyanic acid and acrolein may form as by-products.

ExamplesCatalyst Preparation

Comparative Example 1

5 Ammonium heptamolybdate (49.7 g), ammonium metavanadate (9.8 g) and telluric acid (14.9 g) were dissolved in water at 80°C, and this solution was allowed to cool to 65°C. Oxalic acid dihydrate (5.5 g) was dissolved in a 6.5% w/w solution of niobium oxalate in water (289.9 g), and the resulting niobium oxalate solution, having a temperature of 25°C, was added to the first solution, having a
10 temperature of 65°C. The mixture was dried in a rotary evaporator, and then overnight under vacuum. The dried solid was calcined 1 hour at 275°C in air, and then 2 hours at 600°C in argon.

Example 1

15 The catalyst was prepared using the same procedure as Comparative Example 1, except that the first solution was maintained at 80°C and the niobium oxalate solution, having a temperature of 25°C, was added to the first solution, having a temperature of 80°C.

20 Example 2

25 The catalyst was prepared using the same procedure as Comparative Example 1, except that the first solution was heated to 100°C, the niobium oxalate solution was heated to 100°C, and the niobium oxalate solution, having a temperature of 100°C, was added to the first solution, having a temperature of 100°C.

Comparative Example 2

30 The catalyst was prepared using the same procedure as Comparative Example 1, except that the first solution was allowed to cool to 25°C, and the niobium oxalate solution, having a temperature of 25°C, was added to the first solution, having a temperature of 25°C.

Evaluation and Results

Catalysts were evaluated for the partial oxidation of propane to acrylic acid. The results are shown in Table 1.

5

Table 1

Example	Temp. Of Mixing (°C)	Reaction Temp. (°C)	Propane Conv. (%)	Acrylic Acid Selectivity (%)	Acrylic Acid Yield (%)
Comp. Ex. 1	65/25	364	68.4	54.2	37.0
Ex. 1	80/25	361	79.5	62.3	49.6
Ex. 2	100/100	350	74.2	58.0	43.1
Comp. Ex. 2	25/25	380	63.5	57.6	36.6